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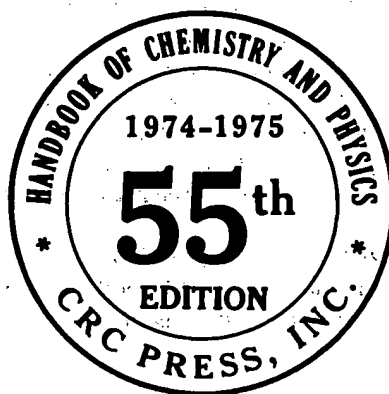
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EXHIBIT I



Handbook OF Chemistry and Physics

A Ready-Reference Book of Chemical and Physical Data



EDITOR

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In collaboration with a large number of professional chemists and physicists whose assistance is acknowledged in the list of general collaborators and in connection with the particular tables or sections involved.

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PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulae	Mol. Wt.	Crystalline form, properties and index of refraction	Density or gravity	Melting point, °C	Solubility, in grams per 100 cc			
							Boiling point, °C	Cold water	Hot water	Other solvents
Sodium										
a289	ferrate (III)	Ferrite, $\text{Na}_2\text{Fe}_2\text{O}_4$	221.67	br. hex pl. or need	4.05					insol. in HCl
a290	ferricyanide	$\text{Na}_3\text{Fe}(\text{CN})_6 \cdot 3\text{H}_2\text{O}$	298.92	red cr, deliq	1.458				67	al. in HCl
a291	ferrocyanate	Yellow prussiate of soda, $\text{Na}_4\text{Fe}(\text{CN})_6 \cdot 10\text{H}_2\text{O}$	484.04	pa yel, monoc.	1.519, 1.530				156.5	al. in HCl
a292	fluoaluminate	Na_3AlF_6	209.94	col, monoc.	2.90	1000			128.6	al. in HCl
a293	fluoantimonate	Na_3SbF_6	258.73	rhomb.	3.375	<1300			128.6	al. in HCl
a294	fluoberyllate	Na_2BeF_4	130.99	wh, rhomb. or monoc	2.475	d 384			1.474	2.94
a295	fluoborate	NaBF_4	109.79	wh, rhomb.	2.475	d 384			108	210
a296	fluoride	Nat. fluorite, NaF	241.99	col, cub. or tet.	2.558	d 993	1095		4.22	al. in HCl
a297	fluoride, hydrogen	$\text{NaF} \cdot \text{H}_2\text{SO}_4$	601.99	col. or wh. cr. powd, rhdr.	2.08				12	57.5
a298	fluoride, orthophosphate	$\text{NaF} \cdot \text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$	422.11		2.2165				12	57.5
a299	fluoroacetate, mono-	$\text{NaC}_2\text{H}_3\text{FO}_2$	100.02	wh. powd.	1.336	200			111	1.4
a300	fluorophosphate, hexa-	$\text{Na}_6\text{P}_2\text{F}_{10}$	185.97		2.369				103.2	al. in HCl
a301	fluorophosphate, mono-	$\text{Na}_2\text{P}_2\text{F}_6$	143.95	col.	2.369	625			25	al. in HCl
a302	fluosilicate	Na_2SiF_6	188.06	col, hex. 1312	2.679	d 312			0.652	2.40
a303	fluosulfate	Na_2SO_4	122.05	shiny leaf, hydr.	1.309	d red heat				al. in HCl
a304	formaldehyde-sulfoxylate	$\text{NaHSO}_3 \cdot \text{CH}_2\text{O} \cdot 2\text{H}_2\text{O}$	154.12	rhomb pr, hydr.	1.309	d 64			97.2	160
a305	formate	NaCHO_2	68.01	col, monoc, deliq.	1.92	d 253			108	al. in HCl
a306	2-furanacrylate	$\text{NaC}_4\text{H}_3\text{O}_5$	160.10	lt brn. powd.	1.919	d			108	al. in HCl
a307	metagermanate	Na_2GeO_3	166.57	wh, monoc, deliq. 1.59	3.31	d 1083			24.6	83
a308	metagermanate, heptahydrate	$\text{Na}_2\text{GeO}_3 \cdot 7\text{H}_2\text{O}$	292.68	col, rhomb.	3.31	d 83			24.6	83
a309	(mono-) d-glutamate	$\text{NaC}_5\text{H}_9\text{NO}_4$	169.11	wh. cr.	1.309	d 80			12	57.5
a310	glycerophosphate, monohydrate	$\text{Na}_2\text{C}_3\text{H}_5\text{O}_6 \cdot \text{H}_2\text{O}$	234.05	yelsh. viol. liq. wh. or powd.	1.474	d 100			14.4	38.9
a311	glycerophosphate, pentahydrate	$\text{Na}_2\text{C}_3\text{H}_5\text{O}_6 \cdot 5\text{H}_2\text{O}$	315.12	wh. pl. ac. or powd.	1.474	d >130			14.4	38.9
a312	gold sulfide	$\text{NaAuS}_2 \cdot 4\text{H}_2\text{O}$	324.08	col, monoc.	1.92	d 80			24.6	83
a313	hydride	NaH	24.00	silver need. 1.470	0.92	d 800			14.4	38.9
a314	hydroxide	NaOH	40.00	wh, deliq. 1.3576	2.130	318.4	1390		347	al. in HCl
a315	iodate	NaIO_3	197.89	wh, rhomb.	4.277	d 300			90	34
a316	metaperiodate	NaIO_4	213.89	col, tet.	4.174	d 300			14.4	38.9
a317	metaperiodate, trihydrate	$\text{NaIO}_4 \cdot 3\text{H}_2\text{O}$	267.94	col, rhomb. deliq.	3.219	d 175			18.78	36.4
a318	paraperiodate	Na_2IO_6	337.85	wh.	3.375	d 800				
a319	(tri-)paraperiodate	$\text{Na}_3\text{H}_2\text{IO}_7$	293.85	col, hexag.	1.605, 1.575	d 80				al. in HCl
a320	iodide	NaI	149.89	col, cub. 1.7745	3.667	d 661	1304		184	302
a321	iodide, dihydrate	$\text{NaI} \cdot 2\text{H}_2\text{O}$	185.92	col, monoc.	2.448	d 752			317.9	1550
a322	iodoplatinate	$\text{Na}_2\text{PtI}_6 \cdot 6\text{H}_2\text{O}$	1110.59	brn, monoc.	3.707	d 50				al. in HCl
a323	iridium chloride	Sodium hexachloroiridate, $\text{Na}_3\text{IrCl}_6 \cdot 12\text{H}_2\text{O}$	690.07	olive cr. rhomb. or trig. rhomb.	1.687	d 50			40	al. in HCl
a324	iron (III) nitrosopentacyanide	$\text{Na}_3[\text{Fe}(\text{CN})_5\text{NO}] \cdot 2\text{H}_2\text{O}$	297.95	ruby red, rhomb.	1.605, 1.575	d 100	160		128.6	al. in HCl
a325	iron (III) oxalate	$\text{Na}_2[\text{Fe}(\text{C}_2\text{O}_4)_3] \cdot 5\text{H}_2\text{O}$	487.96	grn, monoc.	1.973	d 40	300		324	182
a326	iron (III) sulfate	$3\text{Na}_2\text{SO}_4 \cdot \text{Fe}_2(\text{SO}_4)_3 \cdot 6\text{H}_2\text{O}$	934.09	wh, trig. 1.558, 1.613	2.558	d 98	100		140	40
a327	lactate	$\text{NaC}_3\text{H}_5\text{O}_2$	112.06	col or yelsh. liq. very hygro.	1.309	d 17			12	57.5
a328	lithium sulfate	$\text{Na}_2\text{Li}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	376.12	col, ditrig.	2.009	d 50			140	40
a329	magnesium carbonate	$\text{Na}_2\text{CO}_3 \cdot \text{MgCO}_3$	190.31	wh, rhomb.	2.729	d 677	1240/atm			
a330	magnesium sulfate	Nat. bloedite, $\text{Na}_2\text{SO}_4 \cdot \text{MgSO}_4 \cdot 4\text{H}_2\text{O}$	334.48	col, monoc. 1.480, 1.488, 1.489	2.23	d 50				al. in HCl
a331	magnesium tartrate	$\text{Na}_2\text{Mg}(\text{C}_4\text{H}_4\text{O}_6)_2 \cdot 10\text{H}_2\text{O}$	546.59	wh, monoc. pr. or powd.	1.488, 1.489	d 50				al. in HCl